

Screening Stability, Thermochemistry, and Chemical Kinetics of 3-Hydroxybutanoic Acid as a Bifunctional Biodiesel Additive

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Cite This: *J. Phys. Chem. A* 2024, 128, 4068–4082



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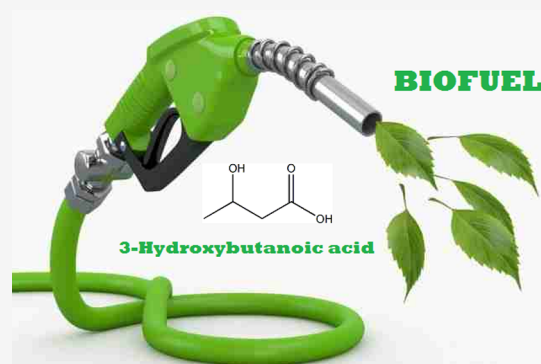


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ABSTRACT: The thermo-kinetic aspects of 3-hydroxybutyric acid (3-HBA) pyrolysis in the gas phase were investigated using density functional theory (DFT), specifically the M06-2X theoretical level in conjunction with the cc-pVTZ basis set. The obtained data were compared with benchmark CBS-QB3 results. The degradation mechanism was divided into 16 pathways, comprising 6 complex fissions and 10 barrierless reactions. Energy profiles were calculated and supplemented with computations of rate coefficients and branching ratios over the temperature range of 600–1700 K at a pressure of 1 bar using transition state theory (TST) and Rice–Ramsperger–Kassel–Marcus (RRKM) methods. Thermodynamics results indicated the presence of six stable conformers within a 4 kcal mol^{−1} energy range. The estimated chemical kinetics results suggested that TST and RRKM approaches are comparable, providing confidence in our calculations. The branching ratio analysis reveals that the dehydration reaction pathway leading to the formation of H₂O and CH₃CH=CHCO₂H dominates entirely at $T \leq 650$ K. At these temperatures, there is a minor contribution from the simple homolytic bond fission reaction, yielding related radicals [CH₃•CHOH + •CH₂CO₂H]. However, at $T \geq 700$ K, this reaction becomes the primary decomposition route. At $T = 1700$ K, there is a minor involvement of a reaction pathway resulting in the formation of CH₃CH(OH)•CH₂ + •CHO(OH) with an approximate contribution of 16%, and a reaction leading to [•CH₃ + •CH₂OHCH₂CO₂H] with around 9%.



1. INTRODUCTION

Biofuel represents a crucial form of renewable energy with the capacity to tackle fundamental worldwide challenges like environmental pollution and energy shortages. Despite biofuels leading to the release of greenhouse gases when consumed, they are considered carbon-neutral fuels, making them environmentally friendly.^{1–3} One specific biofuel variant, biodiesel, consists of alkyl ester fatty acids with long chains (12–20 carbon atoms) derived from biomass, including plants and animals.^{4,5}

Bifunctional organic compounds are recognized as significant substitutes for energy sources due to their multifunctional groups, which can enhance various aspects of their ignition characteristics compared to unifunctional counterparts.^{6,7} One notable bifunctional organic compound is 3-hydroxybutyric acid (3-HBA), possessing both a hydroxyl group (–OH) and a carboxylic group (–COOH). This natural compound is present in human livers through the metabolism of fatty acids⁸ and is found in various organisms, such as the bacteria *Vitis rotundifolia* and *Cupriavidus necator*, among others. 3-HBA holds promise as a precursor for diverse biodegradable plastics, including polyester. In nature, bacteria like *Alcaligenes eutrophus* produce poly(3-hydroxybutyrate) from 3-hydroxybutyric acid.⁹ On a commercial scale, 3-HBA can be derived from poly(3-hydroxybutyrate) through acid hydrolysis.¹⁰

Despite challenges such as experimental shortages and high computational costs, there have been some experimental and computational studies focused on understanding the behavior of real biofuels and biodiesel molecules. These studies have provided valuable insights into the properties and performance of biodiesel.^{6,7,11–33} To comprehend the thermal degradation mechanism of real biodiesel, model biodiesel becomes essential. 3-HBA is regarded as a highly effective model for hydroxycarboxylic acid as a molecular biodiesel additive.

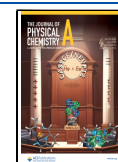
A computational study by Jin-bao et al.³² at the B3LYP/cc-pVTZ theoretical level investigated the decomposition mechanism through the elimination of CO and CO₂ from 2,3,4-hydroxyl-butylaldehyde and 2,3,4-hydroxybutyric acid. Thermo-kinetic parameters were estimated for all pathways at various temperatures. The outcomes revealed six complex fission reactions (three for each compound), with the decarbonylation (CO elimination) from 2,3,4-hydroxybutylaldehyde being exothermic, while the decarboxylation (CO₂

Received: March 1, 2024

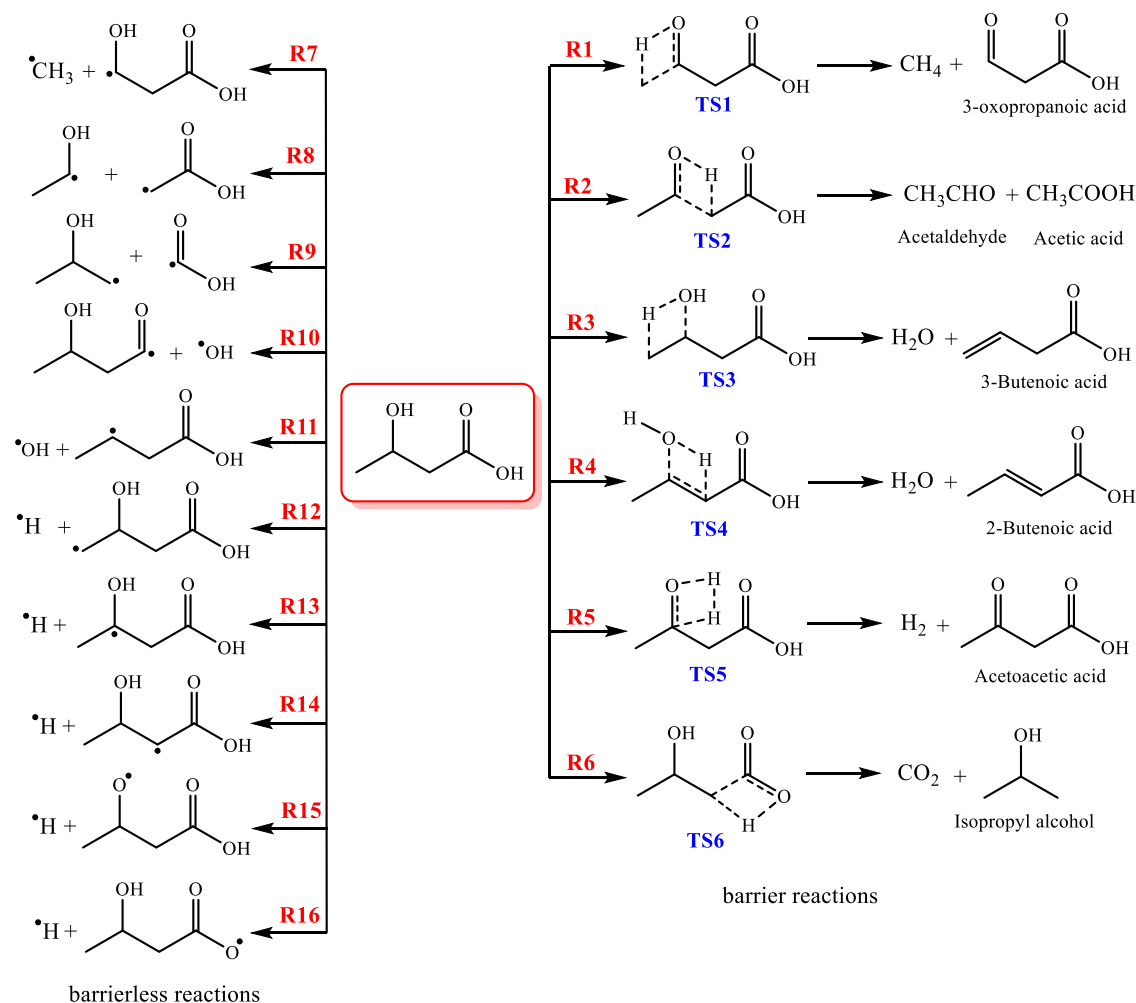
Revised: April 25, 2024

Accepted: May 1, 2024

Published: May 10, 2024



Scheme 1. Possible Decomposition Chemical Channels of 3-HBA



elimination) from 2,3,4-hydroxybutyric acid was endothermic. Notably, the direct activation energy for decarbonyl elimination was much lower than that occurring after dehydration, while for the decarboxyl reaction, the activation energy for decarboxyl elimination after dehydration was much lower than that occurring directly.

The theoretical mechanism of gas-phase pyrolysis of 4-bromobutyric acid to produce butyrolactone and hydrogen bromide was investigated by Tosta et al.³⁴ The authors used both Møller–Plesset perturbation theory of second order (MP2) and density functional theory (DFT) at the PBE/6-31++G(d,p) level of theory to predict the reaction path. Their findings indicated a unimolecular reaction mechanism where the hydroxyl oxygen of the carboxylic group played a role in facilitating bromide removal through nucleophilic substitution. In a separate experimental study, Namysl et al.³⁵ explored the oxidation of butanoic (butyric) and pentanoic (valeric) acids in a jet-stirred reactor under highly diluted conditions at temperatures ranging from 800 to 1100 K and a pressure of 800 Torr. The results revealed a broad spectrum of released products, starting from CO and CO₂ molecules to C5 compounds, including 18 species for butanoic acid and 36 species for pentanoic acid.

To date, there has been no exploration, either computational or experimental, of the pyrolysis of 3-HBA as a bifunctional biodiesel under optimal combustion conditions. To address

this gap, we utilize the M06-2X/cc-pVTZ theoretical level.^{36,37} Subsequently, we compare our calculated reaction energies and energy barriers with high-level composite CBS-QB3 results. The kinetics are assessed using transition state theory (TST)^{38–41} at the high-pressure limit, while the falloff behavior is analyzed statistically analysis employing the Rice–Ramsperger–Kassel–Marcus (RRKM) theory^{42–44} at lower pressures across a temperature range from 600 to 1700 K. Finally, to gain further insights into the pathways studied, we investigate into the results obtained from natural bond orbital (NBO) analysis.^{45,46}

2. COMPUTATIONAL DETAILS

2.1. Potential Energy Surface (PES) Calculations. All quantum chemistry calculations were conducted with the Gaussian 09 suite of programs,⁴⁷ and the molecular structures were visually analyzed with the ChemCraft package.⁴⁸ The geometrical structures and vibrational frequencies of the parent molecule, 3-hydroxybutyric acid, transition states (TSs), and products were optimized using the DFT computational hybrid meta generalized gradient M06-2X functional,³⁷ along with the correlation-consistent polarized valence triplet ζ (cc-pVTZ) basis set.³⁶ To validate the obtained results at the M06-2X/cc-pVTZ theoretical level, a more accurate energy calculation was performed using the multilevel moderate computational cost CBS-QB3 composite method.^{49–51} The approach employed in

this method includes low-level calculations on large basis sets, medium basis sets for second-order Møller–Plesset (MP2) calculations, and small basis sets for high-level correlation corrections [all coordinates are detailed in Table S1 in the Supporting Information]. The five-step CBS-QB3 series of calculations initiates with a geometry optimization at the B3LYP/6-311G(2d,d,p) level, followed by a frequency calculation to acquire thermal corrections, zero-point vibrational energy, and entropic information.^{52–54}

The optimized structures were also confirmed to be real minima by frequency calculations. Frontier molecular orbital (FMO) properties and natural bond orbital (NBO) analysis are measured using the NBO technique. The molecular properties such as electronegativity (χ), chemical potential, ionization potential (IP), chemical hardness (η), softness (ζ), and global electrophilicity index (ψ) were calculated using highest occupied molecular orbital–least unoccupied molecular orbital (HOMO–LUMO) analysis at the same theoretical level using the NBO 5.0 program.⁵⁵

To check the nature of different complex fission TSs, the minimum energy path⁵⁶ was determined using intrinsic reaction coordinate (IRC) calculations^{57,58} at the M06-2X/cc-pVTZ level of theory. The IRC calculations were performed in both directions (forward and reverse) with 20 points, employing a step size of 0.1 amu^{1/2} Bohr.

2.2. Calculation of Absolute Rate Constants. According to Scheme 1, the pyrolysis mechanism of 3-hydroxybutyric acid involves two kinds of chemical pathways: (a) barrier reactions, which include hydrogen atom transfer [R1–R6], and (b) barrierless reactions of simple bond cleavage through reaction pathways R7–R16.

2.2.1. Chemical Kinetic of Barrier Reactions. In the statistical adiabatic channel model (SACM), this frequent observation was attributed to the appearance of adiabatic channel energy barriers, which are a consequence of the interplay of the radial and angular parts of the interaction potential and which, by analogy with centrifugal barriers, contract the available phase space.^{59–63} In the statistical adiabatic channel model, adiabatic channel potential curves $V_i(r)$ are calculated. Their maxima define the channel threshold energies E_{oi} which, for thermal conditions, lead to the “activated complex” partition function^{62–64}

$$Q^* = \sum_i \exp(-E_{oi}/kT) \quad (1)$$

using E_{oi} or Q^* , the usual formalism of statistical rate theory (such as TST) is used. For very low temperatures, only the lowest channel states contribute. The potential curves of these channels can be obtained analytically from perturbation theory such that analytical expressions for channel threshold energies, activated complex partition functions, and capture rate coefficients can be obtained as well.^{62,63} To get accurate chemical kinetics, the KiSThelP program⁶⁵ was employed to compute the rate coefficients of unimolecular barrier reactions [R1–R6], denoted as k_{uni} (in s^{−1}) utilizing transition state theory as follows:

$$k_{\text{uni}} = \kappa_{\text{Eck}} \times \frac{\sigma k_B T}{h} \times \frac{Q_{\text{TS}}}{Q_{3\text{-HBA}}} \times \exp\left(-\frac{[E_{\text{TS}} - E_{3\text{-HBA}}]}{RT}\right) \quad (2)$$

where σ is the reaction pathway degeneracy, $\kappa_{\text{Eck}}(T)$ denotes one-dimensional Eckart correction tunneling,⁶⁶ and k_B and h

are the Boltzmann and Planck constants, respectively. In the above equations, $Q_{3\text{-HBA}}$ and Q_{TS} represent the total molecular partition functions for 3-HBA and the transition state, respectively. The energy corresponding to these functions [$E_{3\text{-HBA}}$ and E_{TS}], including zero-point vibrational contributions, is calculated using Eckert's tunneling correction at different temperatures as follows:

$$\kappa_{\text{Eckart}}(T) = \frac{\exp(\Delta H_f^{\ddagger,0\text{K}}/k_B T)}{k_B T} \int_0^\infty p(E) \exp(-E/k_B T) dE \quad (3)$$

where $\Delta H_f^{\ddagger,0\text{K}}$ represents the zero-point corrected energy barriers in the forward direction and $p(E)$ denotes the probability of transmission through the one-dimensional barrier at energy E . Atmospheric pressures are considered sufficient for reliably calculating the kinetics rate constant using TST. Additionally, the falloff behavior of canonical kinetic rate constants, denoted as $k(T)$, transitioning from the TST limit ($P \rightarrow \infty$) toward the low-pressure limit ($P \rightarrow 0$), is computed using the RRKM theory. The microcanonical kinetics, represented by $k(E)$, are assessed according to unimolecular RRKM theory⁴⁴

$$k(E) = \frac{\sigma N^\ddagger(E)}{h \rho(E)} \quad (4)$$

where $\rho(E)$ represents the vibrational density of states of the reactants and $N^\ddagger(E)$ denotes the total number of states at the transition state. The canonical rate constant $k(T)$ is defined by⁶⁷

$$k(T) = \frac{1}{Q(T)} \int k(E) N(E) \exp(-\beta E) dE \quad (5)$$

where $Q(T)$ represents the internal partition functions of the reactants and β denotes the Boltzmann constant ($\beta = 1/k_B T$). All supplied kinetic data using TST and RRKM theories were obtained using the KiSThelP program. TST provides an upper limit estimate for rate constants in the high-pressure limit,⁶⁷ and RRKM evaluates pressure effects on a microcanonical basis. Collisional stabilization rate constants were computed using Lennard-Jones (L-J) collision rate theory, and the effective collision frequency is given by the following equation:⁶⁸

$$\omega = \beta_c Z_{\text{LJ}}[M] \quad (6)$$

where β_c represents the collisional efficiency, Z_{LJ} is the Lennard-Jones (L-J) collision frequency, and $[M]$ denotes the concentration of the buffer gas.⁶⁹ The retained value for β_c is 0.2.⁶⁵ The collision frequencies (Z_{LJ}) were calculated using the collisional L-J parameters (σ , ϵ/k_B) obtained from the Joback method, which depends on the energy depth (ϵ) of the L-J potential and σ , representing a dimensional scale of the molecular radius.⁷⁰ For helium as a diluent gas, the retained Lennard-Jones potential parameters are $\sigma = 2.64$ Å and $\epsilon/k_B = 10.9$ K, while for 3-HBA, $\sigma = 6.26$ Å and $\epsilon/k_B = 550.91$ K.⁷¹

2.2.2. Chemical Kinetic of Barrierless Reactions. To calculate the rate coefficient of barrierless reactions, the accurate classical method was used. The classical method was tested recently for many comparable systems and gave accurate results.^{6,29–31,72,73}

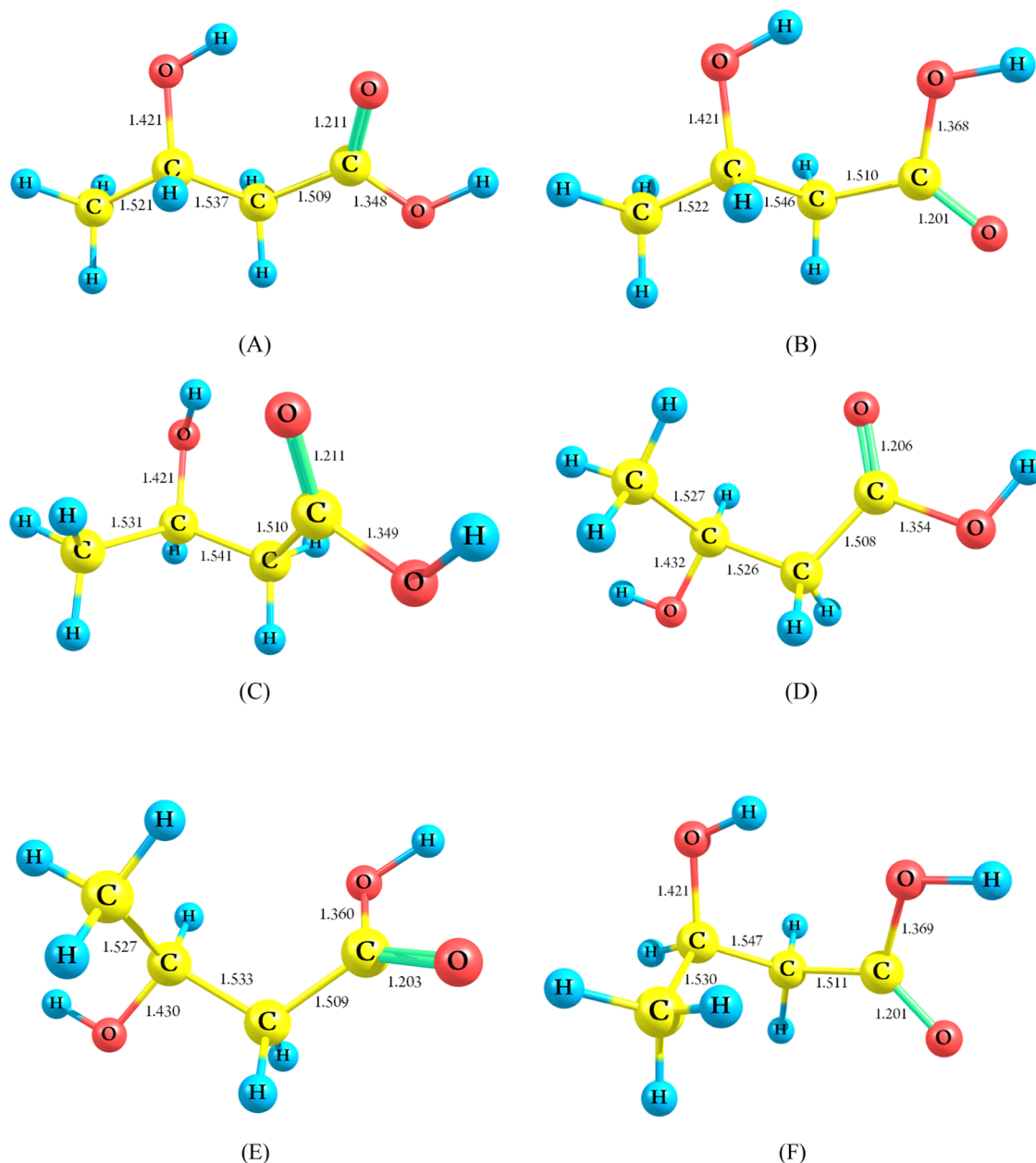


Figure 1. Optimized structures for the different 3-HBA conformers at the CBS-QB3 method.

3. RESULTS AND DISCUSSION

3.1. 3-Hydroxybutyric Acid Conformers. Due to the rotation of the terminal methyl group ($-\text{CH}_3$), the internal hydroxyl group, and the carboxyl group, the 3-HBA molecule exhibits six stable conformers: A, B, C, D, E, and F. The optimized structures of the 3-HBA conformers and their relative energies are depicted in Figures 1 and 2, respectively.

The sum of electronic and zero-point energies as well as the relative energy (ΔE), was investigated in the gas phase, and the results are summarized in Table S2 in the Supporting Information. Upon examination of these conformers, it is evident from the relative energies that conformer A is the most stable, while conformer E is the least stable, with energies of 3.22 (3.24) kcal mol⁻¹, respectively, at the M06-2X and CBS-QB3 (in parentheses) levels relative to conformer A (see Figure 2 and Table S2 in the Supporting Information). Conformers B, C, D, E, and F exhibit energies of 2.62 (2.25),

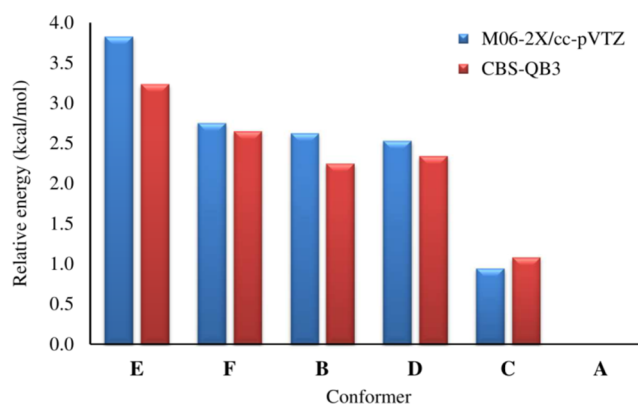


Figure 2. Relative energies of 3-HBA conformers at the M06-2X/cc-pVTZ and CBS-QB3 methods.

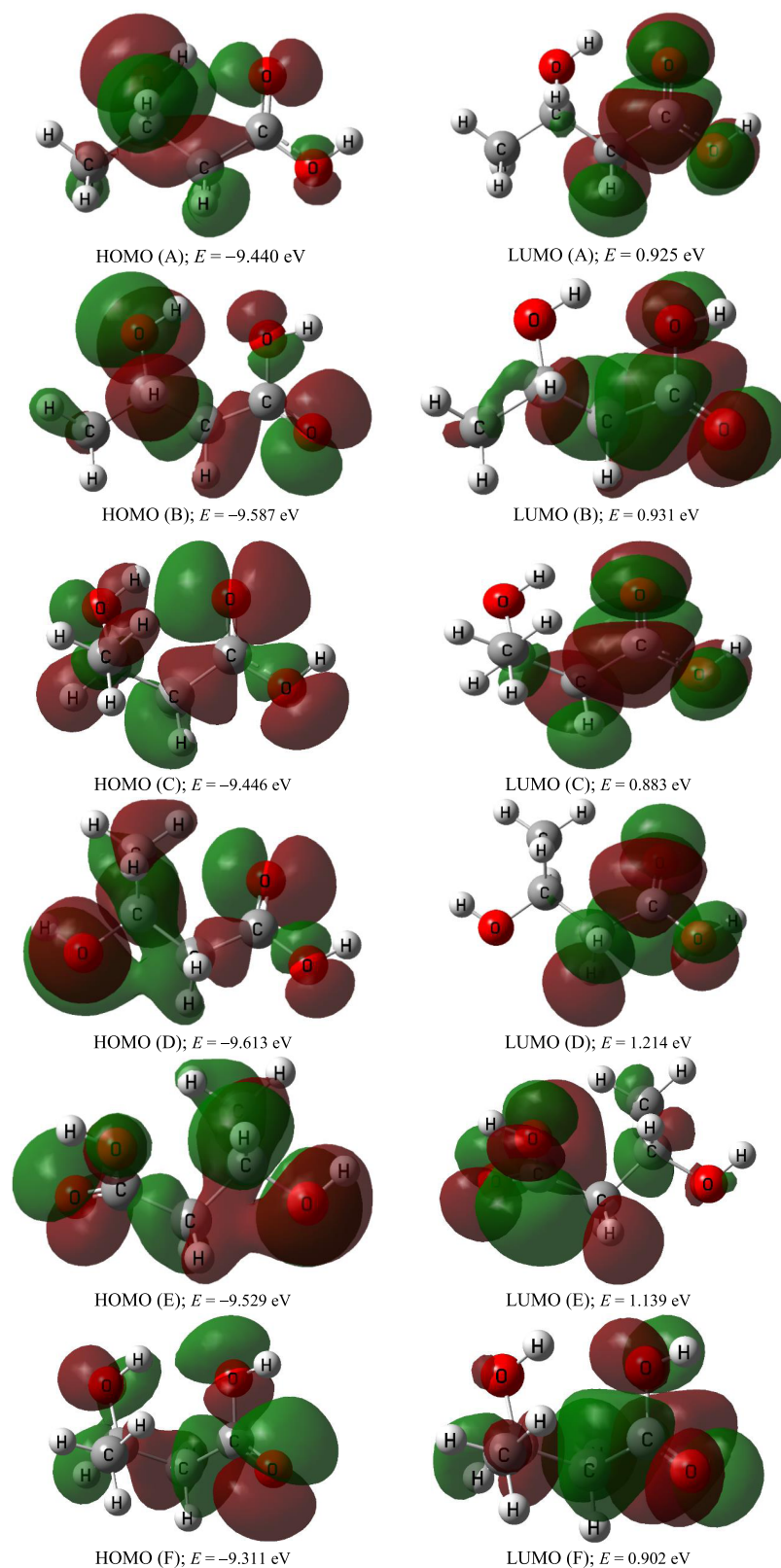


Figure 3. Representative molecular structures of the HOMO and LUMO orbitals of 3-HBA conformers calculated at the M06-2X/cc-pVTZ level of theory.

0.95 (1.08), 2.53 (2.34), 3.83 (3.24), and 2.75 (2.65) kcal mol⁻¹, respectively, relative to conformer A.

3.2. Frontier Molecular Orbital Analysis. Frontier molecular orbital (FMO) energies play a crucial role in

determining the stability, reactivity, optical, and electrical properties of organic molecules. The HOMO–LUMO energy gap explains the concluding charge transfer interaction within the molecule and is useful in determining molecular electrical

transport properties.^{74,75} The estimated LUMO and HOMO of the studied conformers give extensive explanations of their molecular electronic properties. The HOMOs and LUMOs plots at the M06-2X/cc-pVTZ level of theory are presented in Figure 3, while the corresponding HOMO–LUMO energy gaps at the studied methods are shown in Figure 4. The

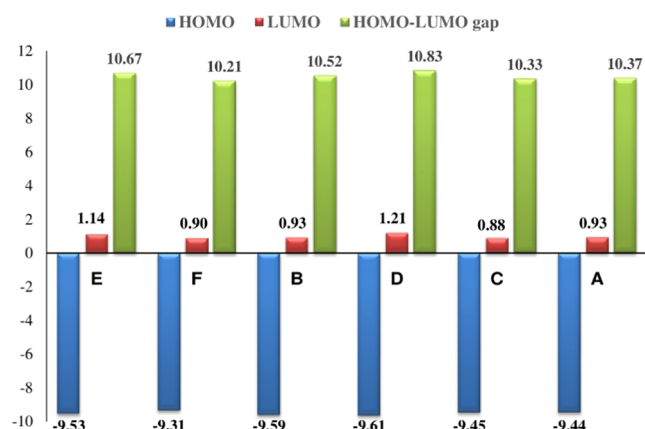


Figure 4. Energy levels of HOMO and LUMO, as well as the HOMO–LUMO energy gaps (in eV) of 3-HBA conformers at the M06-2X/cc-pVTZ level of theory.

HOMO–LUMOs results indicate gaps of 10.37, 10.52, 10.33, 10.83, 10.67, and 10.21 eV for A, B, C, D, E, and F conformers, respectively. Based on the relative energy results, these conformers can be arranged as follows: E > F > B > D > C > A, while the stability order concerning HOMO–LUMOs energy gap is F < C < A < B < E < D.

This study calculates HOMO and LUMO energies using the same theoretical level, focusing on chemical hardness and polarizability. Analysis of the table reveals that conformer D ($\Delta E = 10.83$ eV) is characterized as hard and more stable, indicating lower chemical reactivity. Conversely, conformer F ($\Delta E = 10.21$ eV) is identified as soft and the least stable in the gas phase, signifying higher chemical reactivity.

3.3. Validity of the Studied DFT Method. Figure 5 illustrates the optimized geometries of 3-HBA and its different transition states at the CBS-QB3 method. Meanwhile, Table S3 in the Supporting Information compiles the geometrical parameters of main bond lengths and angles obtained through the CBS-QB3 method. The correlation between the M06-2X/cc-pVTZ and CBS-QB3 results is depicted in Figure 6.

To assess the validity of the employed computational methods, the mean absolute error (MAE), mean signed error (MSE), and root-mean-square error (RMSE) have been computed. The expressions for MSE, MAE, and RMSE are as follows:

$$\text{MAE} = \frac{\sum |\theta_i - E_i|}{n} \quad (7)$$

$$\text{MSE} = \frac{\sum (\theta_i - E_i)}{n} \quad (8)$$

$$\text{RMSE} = \sqrt{\frac{\sum (\theta_i - E_i)^2}{n}} \quad (9)$$

where n represents the total number of observations, and θ_i and E_i denote the M06-2X/cc-pVTZ and CBS-QB3 results, respectively. Utilizing the previously mentioned equations for

MSE, MAE, and RMSE calculations, the findings indicate the following discrepancies in bond lengths: 0.009, 0.021, and 0.026 Å for the C₁–C₂ bond; 0.083, 0.087, and 0.231 Å for the C₂–C₃ bond; 0.016, 0.025, and 0.046 Å for the C₃–C₄ bond; 0.023, 0.025, and 0.043 Å for the C₁–O₂ bond; and 0.029, 0.033, and 0.052 Å for the C₂–O₃ bond. Regarding bond angles, the variations are as follows: −1.82, 5.22, and 10.13° for the C₁–C₂–C₃ angle; −4.30, 4.30, and 9.40° for the C₂–C₃–C₄ angle; −2.13, 2.27, and 5.40° for O₃–C₃–C₂ angle; and −0.70, 0.80, and 1.21° for O₃–C₃–C₄ angle. These results demonstrate a substantial agreement between the CBS-QB3 and M06-2X/cc-pVTZ methods.

3.4. Thermochemistry of Complex Bond Fission Reactions. Complex fission (barrier) reactions refer to reactions that occur through hydrogen transfer or molecular elimination. According to Scheme 1, the 3-hydroxybutyric acid molecule can undergo pyrolysis through six complex fission reactions (R1–R6). The optimized structures of 3-HBA and the resulting transition states (TS1–TS6) are illustrated in Figure 5 and Table S3 in the Supporting Information, as well as with Cartesian structures in Table S1 in the Supporting Information. According to this figure, reactions R1 and R2 exhibit energy barrier reactions involving 1,3-H-transfer reactions through transition states TS1 and TS2, resulting in the production of “methane and 3-oxo-propionic acid” and “acetaldehyde and acetic acid”, respectively. Dehydration (water elimination) of the parent molecule leads to the formation of 3-butenic acid and 2-butenic acid in reactions R3 and R4, respectively. Additionally, acetoacetic acid can be obtained in reaction R5 through the elimination of a hydrogen molecule via transition state TS5. Molecular elimination of CO₂ occurs in reaction R6 yielding isopropyl alcohol. Tables 1 and 2 provide a summary of the barrier and reaction energetic and thermodynamic parameters during the pyrolysis of 3-hydroxybutyric acid using the M06-2X and CBS-QB3 methods, and the corresponding potential energy diagram is presented in Figure 7. The variations in bond lengths and bond angles along the reaction coordinate for the generation of different products via reaction pathways R1–R6 are illustrated in Figure S1 in the SI, utilizing both the M06-2X/cc-pVTZ and CBS-QB3 methods.

Inspection of the B3LYP/6-311G(2d,d,p) geometries obtained for the transition state TS1 along reaction pathway 1 (R1) reveals that the hydrogen atom from the hydroxyl group migrates to the terminal C₄ of the CH₃ (methyl) group, resulting in the production of methane and 3-oxopropionic acid. The reaction requires barrier heights of 90.55 kcal mol^{−1} and reaction energy of 9.68 kcal mol^{−1}, as determined by the CBS-QB3 method. As depicted in Figure 5, investigation of the TS1 structure shows that the elongation of the O₃–H₈ and C₃–C₄ bonds is evident, with increases of 0.110 Å (11.38%) and 0.920 Å (60.49%), respectively, compared to the equilibrium structure computed for 3-HBA. In contrast, the forming C₄–H₈ bond has a larger length than in the isolated methane molecule, and the C₃=O₃ bond length is shortened by 0.119 Å (8.37%).

Reaction 2 leads to the cleavage of the C₂–C₃ and O₃–H₈ bonds through the four-membered ring transition state TS2 to produce acetaldehyde and acetic acid located at 9.14 kcal mol^{−1} above the 3-HBA at the CBS-QB3 method. TS2 results from a simple elongation of the breaking O₃–H₈ and C₂–C₃ bond lengths and the simultaneous shrinkage of the C₃–O₃ distance as a result of the forming of the C₂–H₈ simple bond.

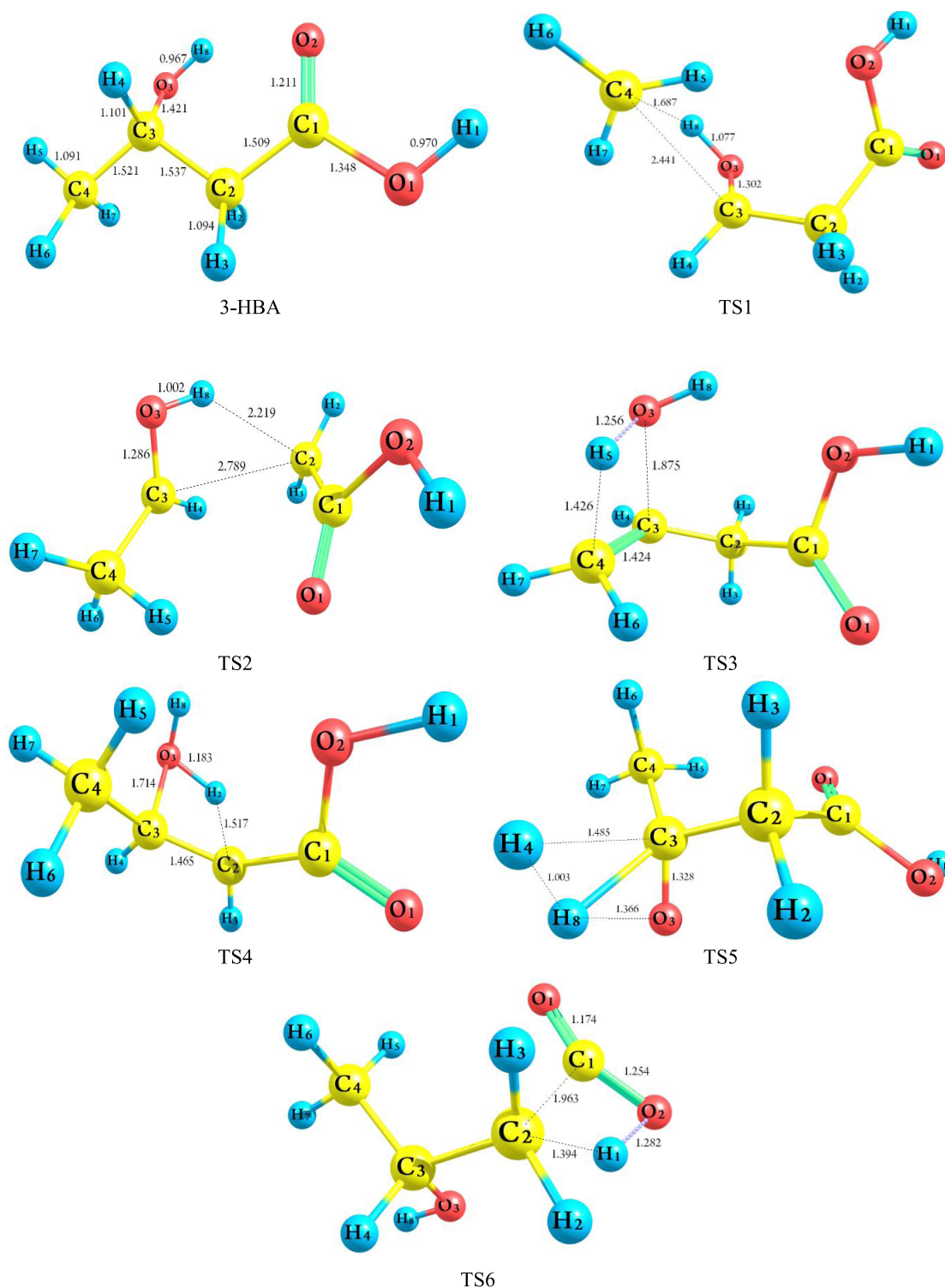


Figure 5. Optimized structure of the transition states at the CBS-QB3 method for unimolecular complex bond fission of 3-HBA [bond lengths are given in angstrom (Å)].

The C_2-C_3 and O_3-H_8 bonds are elongated by 2.789 and 1.002 Å, respectively, and the forming C_2-H_8 bond is shrunk by 2.219 Å. As depicted in Figure 5, the bond formation of the C_2-H_8 and $C_3=O_3$ is reduced by 1.127 Å (50.79%) and 0.082 Å (6.81%), respectively, while the bond breaking of the O_3-H_8 and C_2-C_3 undergoes stretching by 0.035 Å (3.62%) and 1.252 Å (81.46%), respectively.

The transition state TS3 structure obtained along pathway 3 demonstrates that the oxygen atom (O_3) of the hydroxyl group attracts the H_5 atom of the terminal C_4 atom (CH_3 group). Conversely, in the case of the TS4 structure, the O_3 atom captures the hydrogen atom of the internal C_2 atom of the CH_2 group. Among all of the barrier reactions considered in this work, the elimination of a water molecule during chemical reaction R4 is the most favorable kinetic route. The four-

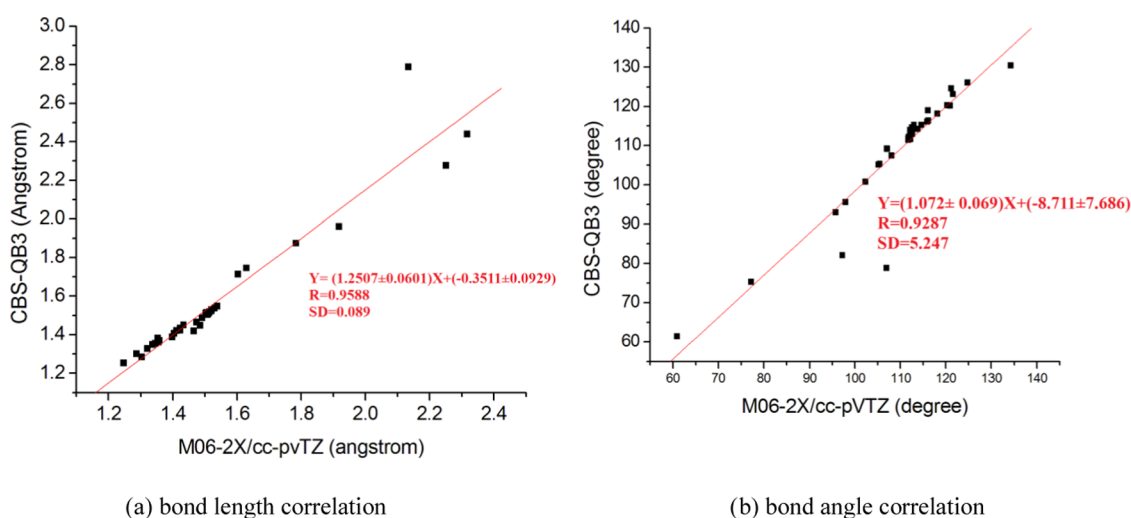


Figure 6. Correlation between CBS-QB3 and M06-2X/cc-pVTZ results: (a) bond lengths and (b) bond angles.

Table 1. Activation Parameters (Energy, Enthalpy, and Gibbs Free Energy) (in kcal mol⁻¹) for Barrier Reactions Encounter 3-HBA Pyrolysis along with Pathways 1–6 at the Studied Methods ($P = 1$ atm)

method	M06-2X/cc-pVTZ			CBS-QB3		
species	$\Delta E_{0\text{K}}^{\ddagger}$	$\Delta H_{298\text{K}}^{\ddagger}$	$\Delta G_{298\text{K}}^{\ddagger}$	$\Delta E_{0\text{K}}^{\ddagger}$	$\Delta H_{298\text{K}}^{\ddagger}$	$\Delta G_{298\text{K}}^{\ddagger}$
3-HBA	0.00	0.00	0.00	0.00	0.00	0.00
3-HBA → TS1	95.27	95.29	95.21	90.55	90.74	90.25
imaginary frequency TS1 (cm ⁻¹)	1329.1i			968.5i		
3-HBA → TS2	80.54	80.89	79.66	77.04	77.39	76.30
imaginary frequency TS2 (cm ⁻¹)	2224.7i			1833.2i		
3-HBA → TS3	67.45	67.45	67.50	68.60	68.64	68.57
imaginary frequency TS3 (cm ⁻¹)	1919.8i			2025.6i		
3-HBA → TS4	58.24	58.10	58.70	60.06	60.16	59.78
imaginary frequency TS4 (cm ⁻¹)	1226.3i			1684.2i		
3-HBA → TS5	86.27	86.28	85.91	85.99	85.99	85.71
imaginary frequency TS5 (cm ⁻¹)	2246.9i			2218.3i		
3-HBA → TS6	71.58	71.87	71.11	72.81	73.11	72.40
imaginary frequency TS6 (cm ⁻¹)	1848.2i			2032.1i		

Table 2. Reaction Parameters (Energy, Enthalpy, and Gibbs Free Energy) (in kcal mol⁻¹) of the 3-HBA Pyrolysis along with Pathways 1–6 at the Studied Methods ($P = 1$ atm)

method	M06-2X/cc-pVTZ			CBS-QB3		
species	$\Delta E_{0\text{K}}$	$\Delta H_{298\text{K}}^{\circ}$	$\Delta G_{298\text{K}}^{\circ}$	$\Delta E_{0\text{K}}$	$\Delta H_{298\text{K}}^{\circ}$	$\Delta G_{298\text{K}}^{\circ}$
3-HBA	0.00	0.00	0.00	0.00	0.00	0.00
3-HBA → pre-RC1	10.41	12.26	7.07			
3-HBA → CH ₄ + CHOCH ₂ CO ₂ H	11.50	12.93	0.25	9.68	11.09	-1.52
3-HBA → pre-RC2	5.55	7.22	1.54			
3-HBA → CH ₃ CHO + CH ₃ CO ₂ H	10.21	10.82	-1.64	9.14	10.18	-3.26
3-HBA → pre-RC3	12.09	13.50	9.89			
3-HBA → H ₂ O + CH ₂ =CHCH ₂ CO ₂ H	17.22	18.83	6.79	14.20	15.81	3.82
3-HBA → pre-RC4	11.78	13.54	8.69			
3-HBA → H ₂ O + CH ₃ CH=CHCO ₂ H	14.32	15.96	4.10	11.25	12.90	0.96
3-HBA → pre-RC5	16.86	18.32	14.08			
3-HBA → H ₂ + CH ₃ CHOCH ₂ CO ₂ H	15.30	17.39	7.10	13.16	15.21	4.80
3-HBA → pre-RC6	-6.59	-5.43	-9.14			
3-HBA → CO ₂ + CH ₃ CHOHCH ₃	-5.35	-4.57	-15.73	-7.26	-6.46	-17.55

membered ring TS4 structure, with an activation energy of 60.06 kcal mol⁻¹, is energetically more favorable than the TS3 structure by 8.54 kcal mol⁻¹ and is less endothermic by 2.95 kcal mol⁻¹, according to the CBS-QB3 method. Inspection of Figure 5 shows that during the progress of reaction R3, the

forming O₃–H₅ bond has a larger length than in the isolated H₂O molecule in the TS3 structure (0.294 Å), and the formation of the C₃=C₄ bond is shortened by 0.097 Å (6.38%). On the contrary, the breaking C₄–H₅ and C₃–O₃ bonds are elongated by 0.335 Å (30.71%) and 0.454 Å

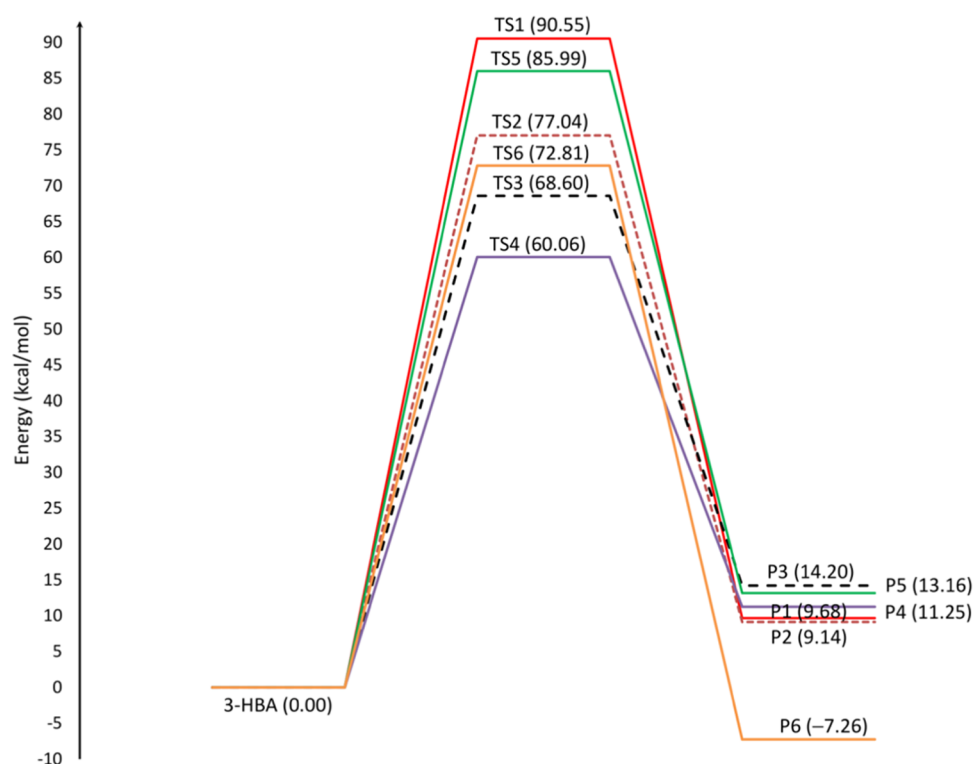


Figure 7. Potential energy profile for barrier reactions involved in the unimolecular degradation mechanism of 3-HBA ($\Delta E_{298\text{ K}}$, $\Delta E_{298\text{ K}}^\ddagger$, in kcal mol⁻¹) at the CBS-QB3 method ($T = 298\text{ K}$, $P = 1\text{ bar}$).

Table 3. Relative Thermodynamic Parameters (in kcal mol⁻¹) as well as Reaction Entropies (in cal mol⁻¹ K⁻¹) for Main Barrier Reactions at Different Temperatures at the CBS-QB3 Method

$T\text{ (K)}$	R3			R4			R6		
	$\Delta H_{298\text{ K}}^\circ$	$\Delta G_{298\text{ K}}^\circ$	$\Delta S_{298\text{ K}}^\circ$	$\Delta H_{298\text{ K}}^\circ$	$\Delta G_{298\text{ K}}^\circ$	$\Delta S_{298\text{ K}}^\circ$	$\Delta H_{298\text{ K}}^\circ$	$\Delta G_{298\text{ K}}^\circ$	$\Delta S_{298\text{ K}}^\circ$
298	15.77	3.76	40.28	12.86	0.91	40.10	-6.48	-17.58	37.23
400	15.92	-0.37	40.74	13.03	-3.21	40.60	-6.54	-21.37	37.09
600	15.88	-8.52	40.67	12.97	-11.33	40.51	-6.90	-28.72	36.37
700	15.73	-12.58	40.44	12.81	-15.37	40.26	-7.15	-32.34	35.99
800	15.51	-16.61	40.16	12.58	-19.38	39.96	-7.42	-35.92	35.63
900	15.26	-20.61	39.85	12.31	-23.36	39.64	-7.69	-39.47	35.30
1000	14.96	-24.58	39.54	12.01	-27.31	39.32	-7.98	-42.98	35.00
1100	14.64	-28.52	39.24	11.68	-31.23	39.00	-8.26	-46.47	34.73
1200	14.30	-32.43	38.94	11.33	-35.11	38.70	-8.55	-49.93	34.48
1300	13.94	-36.31	38.65	10.97	-38.97	38.41	-8.83	-53.37	34.26
1400	13.57	-40.16	38.38	10.59	-42.79	38.13	-9.12	-56.78	34.04
1500	13.19	-43.98	38.12	10.21	-46.59	37.87	-9.40	-60.18	33.83
1600	12.80	-47.78	37.87	9.82	-50.37	37.62	-9.69	-63.55	33.66
1700	12.41	-51.56	37.63	9.42	-54.12	37.38	-9.97	-66.91	33.97

(31.95%), respectively. Furthermore, in the TS4 structure, the bond lengths of O₃-H₂ and C₂=C₃ are shortened by 1.183 Å and 1.465 Å, respectively, while the bond breaking of C₂-H₂ and C₃-O₃ is elongated by 0.423 Å (38.67%) and 0.293 Å (20.62%), respectively.

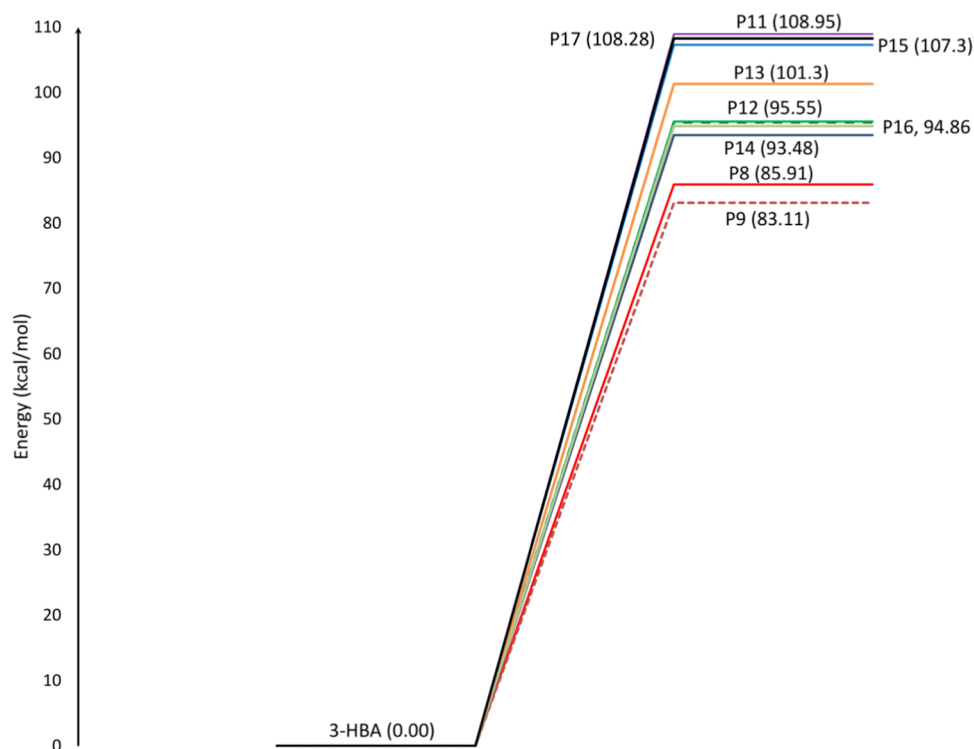
In pathway S, acetoacetic acid can be obtained through the four-membered ring transition state TS5 when the hydrogen atom of the hydroxyl group joins with the hydrogen atom of the tertiary C₃ atom. The reaction is achieved with a barrier height of 85.99 kcal mol⁻¹ and a reaction energy of 13.16 kcal mol⁻¹, according to the CBS-QB3 method. Inspection of the TS5 structure reveals that the bond breaking of O₃-H₈ and C₃-H₄ is stretched by 0.399 Å (41.26%) and 0.384 Å (34.88%), respectively. In comparison, the forming H₄-H₈

bond with a bond distance of 1.003 Å is longer than the hydrogen molecule (0.744 Å) in the TS5 structure, and the formation of the C₃=O₃ bond is shortened by 0.093 Å (6.55%). In other words, the O₃-H₈ and C₃-H₄ bond distances increase, showing the breaking of these bonds (0.967–1.366 Å) and (1.094–1.485 Å) in the related TSs, respectively. The C₃-C₃ bond distance reveals changes from single to double bond character (1.421–1.328 Å) in TSs. The H₄-H₈ bond is forming as the distance between these atoms decreases in the TS (1.003 Å).

As seen from Scheme 1 and Table 1, the removal of the CO₂ molecule from the 3-HBA molecule occurs through the four-membered rings TS6 with an imaginary frequency of 2032.1i at the CBS-QB3 method. In this chemical reaction, the hydrogen

Table 4. Energetic and Thermodynamics Parameters for the Considered Barrierless Reactions Encounter 3-HBA Decomposition (in kcal mol⁻¹) at the Studied Methods (*P* = 1 atm)

method	M06-2X/cc-pVTZ			CBS-QB3		
reaction	$\Delta E_{0\text{ K}}$	$\Delta H^\circ_{298\text{ K}}$	$\Delta G^\circ_{298\text{ K}}$	$\Delta E_{0\text{ K}}$	$\Delta H^\circ_{298\text{ K}}$	$\Delta G^\circ_{298\text{ K}}$
3-HBA	0.00	0.00	0.00	0.00	0.00	0.00
R7: 3-HBA $\rightarrow \bullet\text{CH}_3 + \bullet\text{CH}(\text{OH})\text{CH}_2\text{CO}_2\text{H}$	84.47	86.29	72.65	85.91	87.60	74.33
R8: 3-HBA $\rightarrow \text{CH}_3\bullet\text{CH}(\text{OH}) + \bullet\text{CH}_2\text{CO}_2\text{H}$	81.28	82.48	68.26	83.11	84.25	70.27
R9: 3-HBA $\rightarrow \text{CH}_3\text{CH}(\text{OH})\bullet\text{CH}_2 + \bullet\text{CO}_2\text{H}$	94.45	95.75	81.66	95.47	96.80	82.67
R10: 3-HBA $\rightarrow \bullet\text{OH} + \text{CH}_3\text{CH}(\text{OH})\text{CH}_2\bullet\text{CO}$	109.64	111.37	99.25	108.95	110.65	98.74
R11: 3-HBA $\rightarrow \bullet\text{OH} + \text{CH}_3\text{CH}\bullet\text{CH}_2\text{CO}_2\text{H}$	95.18	97.06	84.17	95.55	97.43	84.51
R12: 3-HBA $\rightarrow \bullet\text{H} + \bullet\text{CH}_2\text{CH}(\text{OH})\text{CH}_2\text{CO}_2\text{H}$	99.20	100.83	92.02	101.30	102.96	94.09
R13: 3-HBA $\rightarrow \bullet\text{H} + \text{CH}_3\text{C}\bullet(\text{OH})\text{CH}_2\text{CO}_2\text{H}$	92.18	93.92	84.37	93.48	95.25	85.59
R14: 3-HBA $\rightarrow \bullet\text{H} + \text{CH}_3\text{CH}(\text{OH})\bullet\text{CHCO}_2\text{H}$	105.14	106.61	97.67	107.30	108.82	98.52
R15: 3-HBA $\rightarrow \bullet\text{H} + \text{CH}_3\text{CH}(\text{O}\bullet)\text{CH}_2\text{CO}_2\text{H}$	93.73	95.38	86.50	94.86	96.37	87.94
R16: 3-HBA $\rightarrow \bullet\text{H} + \text{CH}_3\text{CH}(\text{OH})\text{CH}_2\text{COO}\bullet$	109.27	110.83	101.89	108.28	109.79	100.81

**Figure 8.** Potential energy profiles ($\Delta E_{298\text{ K}}$ and $\Delta E_{298\text{ K}}^\ddagger$ are in kcal mol⁻¹) for barrierless reactions involved in the unimolecular reactions of 3-HBA at the CBS-QB3 method (*T* = 298 K, *P* = 1 bar).

atom (H₁) on the hydroxyl group on C₁ is transferred to the C₂ atom, and simultaneously, the C₁–C₂ bond is broken to produce CO₂ (see Figure 5). The direct removal of the carboxyl group to produce carbon dioxide and isopropyl alcohol is the only exothermic reaction among all investigated channels [R1–R6]. Through the elimination of CO₂, the formed double bond C₁–O₂ as well as the single bond C₂–H₁ is shortened by 0.094 Å (8.10%) and 0.19 Å (17.40%), respectively, while the broken C₁–C₂ and O₂–H₁ bonds stretch by 0.454 Å (30.09%) and 0.312 Å (32.17%), respectively.

The parameter *n_T* describes the position of the TS structure along the reaction coordinate⁷⁶

$$n_T = \frac{1}{2 - (\Delta G_r / \Delta G^\ddagger)} \quad (10)$$

According to this equation, when *n_T* < 0.5 (indicating an early TS), the TS structure is similar to the reactant; conversely, when *n_T* > 0.5, it resembles the product (suggesting a late TS).³³ In the pyrolysis of 3-HBA through pathways 1–6, the TS structures are identical to the 3-HBA and the associated products (P1–P6). By optimizing all stationary points along with reactions 1–6 and determining the activation and reaction Gibbs free energy, the *n_T* values for pathways 1–6 are approximately 0.52, 0.54, 0.44, 0.48, 0.44, and 0.74, respectively. This indicates that reactions R3 and R5, with an *n_T* value of 0.44, closely resemble 3-HBA, while reaction R6 resembles the related product (CO₂ and CH₃CHOHCH₃).

Table 3 provides a summary of the change in standard thermodynamic parameters ΔG° , ΔH° , and ΔS° for the most favorable barrier reactions exhibiting lower energy barrier heights within the studied temperature ranges, as determined by the CBS-QB3 method. The results indicate that all

Table 5. TST and RRKM (in Parentheses) Rate Constants (in s^{-1}) for 3-HBA Pyrolysis Via Barrier Reactions over the Temperature Range 600–1700 K at the CBS-QB3 Method ($P = 1 \text{ bar}$)

$T \text{ (K)}$	k_{R1}	k_{R2}	k_{R3}	k_{R4}	k_{R5}	k_{R6}
600	3.52×10^{-20} (2.83×10^{-20})	5.87×10^{-15} (5.72×10^{-15})	1.53×10^{-11} (4.70×10^{-12})	1.32×10^{-8} (6.19×10^{-9})	4.42×10^{-18} (9.60×10^{-19})	3.50×10^{-13} (1.05×10^{-13})
650	1.31×10^{-17} (1.09×10^{-17})	9.51×10^{-13} (9.27×10^{-13})	1.14×10^{-9} (4.34×10^{-10})	6.16×10^{-7} (3.28×10^{-7})	9.09×10^{-16} (2.66×10^{-16})	3.43×10^{-11} (1.28×10^{-11})
700	2.09×10^{-15} (1.79×10^{-15})	7.42×10^{-11} (7.30×10^{-11})	4.77×10^{-8} (2.21×10^{-8})	1.69×10^{-5} (9.94×10^{-6})	9.16×10^{-14} (3.32×10^{-14})	1.81×10^{-9} (7.93×10^{-10})
750	1.72×10^{-13} (1.50×10^{-13})	3.28×10^{-9} (3.22×10^{-9})	1.23×10^{-6} (6.18×10^{-7})	3.04×10^{-4} (1.92×10^{-4})	5.16×10^{-12} (2.19×10^{-12})	5.73×10^{-8} (2.84×10^{-8})
800	8.20×10^{-12} (7.27×10^{-12})	9.06×10^{-8} (8.90×10^{-8})	2.15×10^{-5} (1.19×10^{-5})	3.84×10^{-3} (2.57×10^{-3})	1.80×10^{-10} (8.59×10^{-11})	1.20×10^{-6} (6.53×10^{-7})
850	2.48×10^{-10} (2.24×10^{-10})	1.68×10^{-6} (1.67×10^{-6})	2.75×10^{-4} (1.63×10^{-4})	3.60×10^{-2} (2.54×10^{-2})	4.17×10^{-9} (2.19×10^{-9})	1.76×10^{-5} (1.04×10^{-5})
900	5.19×10^{-9} (4.73×10^{-9})	2.28×10^{-5} (2.26×10^{-5})	2.64×10^{-3} (1.67×10^{-3})	2.68×10^{-1} (1.96×10^{-1})	6.92×10^{-8} (3.93×10^{-8})	1.95×10^{-4} (1.23×10^{-4})
950	7.90×10^{-8} (7.25×10^{-8})	2.37×10^{-4} (2.34×10^{-4})	2.02×10^{-2} (1.34×10^{-2})	1.61×10^0 (1.22×10^0)	8.60×10^{-7} (5.21×10^{-7})	1.70×10^{-3} (1.12×10^{-3})
1000	9.17×10^{-7} (8.49×10^{-7})	1.94×10^{-3} (1.92×10^{-3})	1.27×10^{-1} (8.77×10^{-2})	8.11×10^0 (6.33×10^0)	8.36×10^{-6} (5.34×10^{-6})	1.18×10^{-2} (8.18×10^{-3})
1050	8.43×10^{-6} (7.88×10^{-6})	1.29×10^{-2} (1.29×10^{-2})	6.73×10^{-1} (4.81×10^{-1})	3.53×10^1 (2.82×10^1)	6.58×10^{-5} (4.40×10^{-5})	6.92×10^{-2} (4.96×10^{-2})
1100	6.35×10^{-5} (5.98×10^{-5})	7.30×10^{-2} (7.29×10^{-2})	3.05×10^0 (2.27×10^0)	1.35×10^2 (1.10×10^2)	4.32×10^{-4} (3.00×10^{-4})	3.44×10^{-1} (2.56×10^{-1})
1150	4.01×10^{-4} (3.81×10^{-4})	3.56×10^{-1} (3.55×10^{-1})	1.23×10^1 (9.34×10^0)	4.56×10^2 (3.79×10^2)	2.41×10^{-3} (1.74×10^{-3})	1.50×10^0 (1.14×10^0)
1200	2.19×10^{-3} (2.08×10^{-3})	1.53×10^0 (1.52×10^0)	4.39×10^1 (3.43×10^1)	1.40×10^3 (1.19×10^3)	1.17×10^{-2} (8.69×10^{-3})	5.84×10^0 (4.53×10^0)
1250	1.04×10^{-2} (9.96×10^{-3})	5.79×10^0 (5.78×10^0)	1.44×10^2 (1.13×10^2)	3.98×10^3 (3.38×10^3)	5.04×10^{-2} (3.82×10^{-2})	2.02×10^1 (1.61×10^1)
1300	4.40×10^{-2} (4.23×10^{-2})	1.99×10^1 (1.99×10^1)	4.25×10^2 (3.43×10^2)	1.03×10^4 (8.91×10^3)	1.94×10^{-1} (1.51×10^{-1})	6.43×10^1 (5.18×10^1)
1350	1.68×10^{-1} (1.61×10^{-1})	6.25×10^1 (6.23×10^1)	1.17×10^3 (9.54×10^2)	2.52×10^4 (2.18×10^4)	6.81×10^{-1} (5.36×10^{-1})	1.87×10^2 (1.53×10^2)
1400	5.77×10^{-1} (5.59×10^{-1})	1.81×10^2 (1.80×10^2)	2.98×10^3 (2.47×10^3)	5.75×10^4 (5.01×10^4)	2.16×10^0 (1.74×10^0)	5.05×10^2 (4.20×10^2)
1450	1.84×10^0 (1.78×10^0)	4.88×10^2 (4.85×10^2)	7.10×10^3 (5.99×10^3)	1.24×10^5 (1.09×10^5)	6.40×10^0 (5.24×10^0)	1.27×10^3 (1.07×10^3)
1500	5.43×10^0 (5.26×10^0)	1.23×10^3 (1.22×10^3)	1.62×10^4 (1.37×10^4)	2.54×10^5 (2.23×10^5)	1.77×10^1 (1.46×10^1)	3.03×10^3 (2.58×10^3)
1550	1.48×10^1 (1.45×10^1)	2.92×10^3 (2.90×10^3)	3.47×10^4 (2.97×10^4)	4.98×10^5 (4.37×10^5)	4.57×10^1 (3.83×10^1)	6.83×10^3 (5.85×10^3)
1600	3.83×10^1 (3.74×10^1)	6.58×10^3 (6.52×10^3)	7.15×10^4 (6.18×10^4)	9.45×10^5 (8.20×10^5)	1.12×10^2 (9.44×10^1)	1.46×10^4 (1.26×10^4)
1650	9.36×10^1 (9.14×10^1)	1.41×10^4 (1.39×10^4)	1.40×10^5 (1.21×10^5)	1.71×10^6 (1.47×10^6)	2.58×10^2 (2.21×10^2)	2.99×10^4 (2.60×10^4)
1700	2.17×10^2 (2.12×10^2)	2.89×10^4 (2.85×10^4)	2.66×10^5 (2.29×10^5)	3.00×10^6 (2.55×10^6)	5.71×10^2 (4.90×10^2)	5.87×10^4 (5.13×10^4)

Table 6. Calculated Rate Constants for Simple Homolytic Bond Fission Reactions [R7–R11] Based on the Computed CBS-QB3 Energies ($P = 1$ bar)

T (K)	k_{R7}	k_{R8}	k_{R9}	k_{R10}	k_{R11}
600	3.09×10^{-11}	1.23×10^{-9}	2.95×10^{-13}	2.11×10^{-20}	6.02×10^{-15}
650	8.06×10^{-9}	2.64×10^{-7}	1.42×10^{-10}	2.54×10^{-17}	3.10×10^{-12}
700	9.41×10^{-7}	2.60×10^{-5}	2.80×10^{-8}	1.10×10^{-14}	6.50×10^{-10}
750	5.77×10^{-5}	1.39×10^{-3}	2.72×10^{-6}	2.12×10^{-12}	6.66×10^{-8}
800	2.10×10^{-3}	4.45×10^{-2}	1.48×10^{-4}	2.10×10^{-10}	3.78×10^{-6}
850	4.96×10^{-2}	9.45×10^{-1}	4.98×10^{-3}	1.21×10^{-8}	1.33×10^{-4}
900	8.20×10^{-1}	1.42×10^1	1.13×10^{-1}	4.39×10^{-7}	3.15×10^{-3}
950	1.00×10^1	1.59×10^2	1.82×10^0	1.09×10^{-5}	5.32×10^{-2}
1000	9.46×10^1	1.39×10^3	2.22×10^1	1.96×10^{-4}	6.72×10^{-1}
1050	7.16×10^2	9.86×10^3	2.13×10^2	2.66×10^{-3}	6.66×10^0
1100	4.48×10^3	5.82×10^4	1.64×10^3	2.83×10^{-2}	5.31×10^1
1150	2.38×10^4	2.92×10^5	1.06×10^4	2.44×10^{-1}	3.55×10^2
1200	1.09×10^5	1.28×10^6	5.83×10^4	1.76×10^0	2.00×10^3
1250	4.43×10^5	4.97×10^6	2.79×10^5	1.08×10^1	9.84×10^3
1300	1.60×10^6	1.73×10^7	2.11×10^6	5.71×10^1	4.27×10^4
1350	5.25×10^6	5.47×10^7	4.44×10^6	2.68×10^2	1.66×10^5
1400	1.57×10^7	1.58×10^8	1.52×10^7	1.12×10^3	5.81×10^5
1450	4.35×10^7	4.26×10^8	4.78×10^7	4.24×10^3	1.87×10^6
1500	1.13×10^8	1.07×10^9	1.38×10^8	2.19×10^4	5.52×10^6
1550	2.71×10^8	2.51×10^9	3.74×10^8	4.63×10^4	1.52×10^7
1600	6.18×10^8	5.61×10^9	9.45×10^8	1.37×10^5	3.93×10^7
1650	1.34×10^9	1.19×10^{10}	2.25×10^9	3.77×10^5	1.49×10^8
1700	2.75×10^9	2.39×10^{10}	5.09×10^9	9.75×10^5	2.21×10^8

investigated pathways exhibit positive entropy, and all thermodynamic parameters show an inverse relationship with Kelvin temperature; in other words, as the temperature increases, all thermodynamic values decrease. As can be seen from Table 3, pathway R6 is identified as a spontaneous ($\Delta G^\circ < 0$) and exothermic ($\Delta H^\circ < 0$) reaction, while pathways R3 and R4 are characterized as spontaneous ($\Delta G^\circ < 0$) and endothermic ($\Delta H^\circ > 0$) reactions.

3.5. Thermochemistry of Simple Bond Fission Reactions. Table 4 presents thermodynamic data for various simple homolytic bond fission reactions (barrierless reactions), calculated using the M06-2X and CBS-QB3 methods. The potential energy profile is depicted in Figure 8. Inspection of different barrierless reactions during the decomposition of 3-HBA indicates that the production of $\bullet\text{CH}_2\text{COOH}$ and $\text{CH}_3\text{C}\bullet\text{HOH}$ occurs through the cleavage of the $\text{C}_2\text{--C}_3$ bond, exhibiting the lowest endothermic energy ($83.1 \text{ kcal mol}^{-1}$) among all simple fission reactions.

The production of hydroxyl radical through reaction R11 can be accomplished via the thermal decomposition of the $\text{C}_3\text{--O}_3$ bond. The reaction requires a reaction energy of 95.18 (95.55) kcal mol^{-1} at the M06-2X (CBS-QB3) method, which is close to the obtained results for 2-butanol (95 kcal mol^{-1}) and 2-methoxyethanol ($97.3 \text{ kcal mol}^{-1}$).^{6,77} The removal of the hydroxyl group via chemical reaction R10 from the acidic COOH is the most costly endothermic reaction, with an energy of 109.64 (108.95) kcal mol^{-1} at the M06-2X (CBS-QB3) method. Due to the energy overlap between complex and simple bond cleavage reactions, simple homolytic bond fission reactions may compete with complex ones at higher temperatures.

3.6. Chemical Kinetic Simulations. Table 5 presents the rate constants for all barrier reactions [R1–R6] using the TST and RRKM theories incorporating the Eckert tunneling coefficient. These kinetic rate constants are calculated at a

pressure of 1 bar over the temperature ranging from 600 to 1700 K. Table 6 provides the rate constants for simple homolytic bond fission reactions [R7–R11] under the same conditions.

The obtained results in Table 5 demonstrate a good agreement between the TST and RRKM rate constants, with individual rate constants increasing as the temperature rises, indicating a positive temperature dependency. The accompanied Eckert tunneling corrections during TST calculations are given in Table S4 in the SI. These results suggest that tunneling is effective for pathways R3–R6 up to 1400 K, while it is considered negligible for other reactions within the applied temperature range.

The branching ratio of the main routes of 3-HBA pyrolysis within a temperature range of 600–1700 K is detailed in Table 7. The outcomes reveal the predominant occurrence of water elimination through reaction pathway R4 [$3\text{-HBA} \rightarrow \text{H}_2\text{O} + \text{CH}_3\text{CH}=\text{CHCO}_2\text{H}$], leading to the formation of 2-butenic acid, with a minor contribution from reaction pathway R8 [$3\text{-HBA} \rightarrow \text{CH}_3\bullet\text{CH}(\text{OH}) + \bullet\text{CH}_2\text{CO}_2\text{H}$] at $T \leq 650$ K. Above 700 K, reaction R8 becomes the primary decomposition route, with a small contribution from reaction pathway R9 [$3\text{-HBA} \rightarrow \text{CH}_3\text{CH}(\text{OH})\bullet\text{CH}_2 + \bullet\text{CO}_2\text{H}$] and reaction R7 [$3\text{-HBA} \rightarrow \bullet\text{CH}_3 + \bullet\text{CH}(\text{OH})\text{CH}_2\text{CO}_2\text{H}$] which are approximately 16 and 9%, respectively, at 1700 K.

4. CONCLUSIONS

The pyrolysis of 3-hydroxybutyric acid (3-HBA) in the gas phase was computationally investigated using density functional theory (M06-2X) in conjunction with the correlation consistent polarized valence triplet ζ (cc-pVTZ) basis set, as well as the CBS-QB3 composite method. Energy profiles were obtained and supplemented with calculations of rate coefficients and branching ratios at a pressure of 1 bar using conventional transition state theory (TST) and statistical

Table 7. Branching Ratio (Γ) of the Main Pathways R3, R4, R7, R8, R9, and R11 in the Pyrolysis Process of 3-Hydroxybutyric Acid

T (K)	R3	R4	R7	R8	R9	R11
600	0.11	91.15	0.21	8.53	0.00	0.00
650	0.13	69.22	0.91	29.72	0.02	0.00
700	0.11	38.46	2.14	59.22	0.06	0.00
750	0.07	17.34	3.29	79.13	0.15	0.00
800	0.04	7.58	4.16	87.91	0.29	0.01
850	0.03	3.48	4.79	91.21	0.48	0.01
900	0.02	1.74	5.33	92.15	0.73	0.02
950	0.01	0.93	5.81	92.16	1.06	0.03
1000	0.01	0.53	6.23	91.72	1.46	0.04
1050	0.01	0.33	6.61	91.03	1.96	0.06
1100	0.00	0.21	6.94	90.22	2.54	0.08
1150	0.00	0.14	7.27	89.24	3.24	0.11
1200	0.00	0.10	7.53	88.21	4.02	0.14
1250	0.00	0.07	7.77	87.09	4.89	0.17
1300	0.00	0.05	7.59	82.13	10.03	0.20
1350	0.00	0.04	8.13	84.69	6.89	0.26
1400	0.00	0.03	8.29	83.33	8.04	0.31
1450	0.00	0.02	8.39	82.03	9.20	0.36
1500	0.00	0.02	8.50	80.62	10.44	0.42
1550	0.00	0.02	8.54	79.18	11.78	0.48
1600	0.00	0.01	8.57	77.76	13.11	0.54
1650	0.00	0.01	8.54	76.08	14.41	0.95
1700	0.00	0.01	8.61	74.78	15.91	0.69

Rice–Ramsperger–Kassel–Marcus (RRKM) theory. This study specifically focused on the analysis of simple and complex bond fission unimolecular reactions at a pressure of 1 bar across a temperature range from 600 to 1700 K. The obtained results can be summarized as follows:

- [1]. At low temperatures, the most kinetically favorable reaction in 3-HBA pyrolysis is the removal of a water molecule through 1,3-hydrogen transfer to produce 2-butenic acid. Meanwhile, the carbon dioxide elimination pathway is more thermodynamically favorable compared to other reactions.
- [2]. At higher temperatures, the main route shifts to the simple homolytic bond fission of C₂–C₃ through reaction R7 [i.e., 3-HBA → •CH₃ + •CH(OH)-CH₂CO₂H].
- [3]. All available chemical reactions involved in the pyrolysis of 3-HBA are endothermic, except for the release of CO₂ molecules [i.e., 3-HBA → CO₂ + CH₃CHOHCH₃].
- [4]. The elimination of the hydroxyl radical of the alcohol is 13.3 kcal mol⁻¹ easier than that of the carboxylic acid.

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/acs.jpca.4c01338>.

The optimized geometry of 3-HBA and transition state structures using the CBS-QB3 method (Table S1); the standard energies and relative energies of all examined conformers of 3-HBA at the M06-2X/cc-pVTZ and CBS-QB3 methods (Table S2); bond lengths (Å) and bond angles (°) for 3-HBA and its TSs using CBS-QB3 and M06-2X/cc-pVTZ (in parentheses) theoretical methods (Table S3); Eckart tunneling corrections for

the studied chemical reactions over the temperature range 600–1700 K at the CBS-QB3 method (Table S4); change of bond lengths along reaction coordinates for the formation of different products at the M06-2X/cc-pVTZ level of theory (Figure S1) (PDF)

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Notes

The authors declare no competing financial interest.

■ ACKNOWLEDGMENTS

M.A.A.-R. thanks the Egyptian government and Suez University for their support and the super computational lab at the Mathematics Department, Faculty of Science, Suez University. A.S. expresses gratitude to the Gdańsk University of Technology for their support under the “Excellence Initiative-Research University” (IDUB) program. H.Z. thanks the Key Laboratory of Marine Power Engineering and Technology of the Ministry of Transport (Wuhan University of Technology, No. KLMPET2022-01) for their support during the study. Some parts of the computations were carried out using the computers of the Center of Informatics at the Tricity Academic Supercomputer & Network. This work was partially supported by the Joint Fund of Ministry of Education 8091B032244 and the Open Fund of the State Key Laboratory of Engines of China K2023-11.

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